Modern atomic Theory Practice Key

1. How does an orbit differ from an orbital?
2. How many electrons can occupy a 6p orbital?
3. What is the complete electron configuration of an atom of silicon?
4. What is the shorthand electron configuration of an atom of cadmium (Cd)?
5. How does atomic size change as you move across the periodic table to the right? What is the explanation for this change?
6. Circle the atom with the higher ionization energy from the following pairs.

Sr or $I$

$K$ or Cs

1. Show the orbital diagram for an atom of nitrogen.
2. An atom of neon is isoelectronic with an Mg+2 ion. What is meant by this statement?